



## INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

<p>(51) International Patent Classification 6 : C01B 3/38, B01J 23/72, H01M 8/06</p>	<p>A1</p>	<p>(11) International Publication Number: <b>WO 96/00186</b> (43) International Publication Date: 4 January 1996 (04.01.96)</p>
<p>(21) International Application Number: PCT/GB95/01500 (22) International Filing Date: 26 June 1995 (26.06.95) (30) Priority Data: 9412786.7 24 June 1994 (24.06.94) GB (71) Applicant (for all designated States except US): JOHNSON MATTHEY PUBLIC LIMITED COMPANY [GB/GB]; 78 Hatton Garden, London EC1N 8JP (GB). (72) Inventors; and (75) Inventors/Applicants (for US only): EDWARDS, Neil [GB/GB]; 6 Holkam Close, Tilehurst, Reading RG3 6BZ (GB). FROST, Jonathon, Charles [GB/GB]; Shiplake Farm Cottage, Chiltern Road, Peppard Common, Henley on Thames RG9 5HY (GB). JONES, Ann-Marie [GB/GB]; 71 Kingsway, Caversham Park Village, Reading RG4 6RH (GB). PIGNON, John, Frederick [GB/GB]; Benson, Oxon OX9 6SN (GB). (74) Agent: WISHART, Ian, Carmichael; Johnson Matthey Technology Centre, Blounts Court, Sonning Common, Reading RG4 9NH (GB).</p>		<p>(81) Designated States: AU, CA, JP, US, European patent (AT, BE, CH, DE, DK, ES, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE).  Published With international search report.</p>
<p>(54) Title: SELF-SUSTAINING HYDROGEN GENERATOR</p> <p>(57) Abstract</p> <p>A self-starting, self-sustaining hydrogen generator comprises a reactor packed with a mass of mixed catalyst, containing supported copper and palladium. A feedstock, of e.g. methanol and an oxygen source such as air, is injected at high velocity into the mass of catalyst through a multipoint entry such as a tube of porous ceramic. The mass of catalyst is preferably configured concentrically around the tube injector, so that the fluid-flow follows a radial path through the reactor. The product gas is high in hydrogen and can be used as a feed for a fuel cell.</p>		

**FOR THE PURPOSES OF INFORMATION ONLY**

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AT	Austria	GB	United Kingdom	MR	Mauritania
AU	Australia	GE	Georgia	MW	Malawi
BB	Barbados	GN	Guinea	NE	Niger
BE	Belgium	GR	Greece	NL	Netherlands
BF	Burkina Faso	HU	Hungary	NO	Norway
BG	Bulgaria	IE	Ireland	NZ	New Zealand
BJ	Benin	IT	Italy	PL	Poland
BR	Brazil	JP	Japan	PT	Portugal
BY	Belarus	KE	Kenya	RO	Romania
CA	Canada	KG	Kyrgyzstan	RU	Russian Federation
CF	Central African Republic	KP	Democratic People's Republic of Korea	SD	Sudan
CG	Congo	KR	Republic of Korea	SE	Sweden
CH	Switzerland	KZ	Kazakhstan	SI	Slovenia
CI	Côte d'Ivoire	LI	Liechtenstein	SK	Slovakia
CM	Cameroon	LK	Sri Lanka	SN	Senegal
CN	China	LU	Luxembourg	TD	Chad
CS	Czechoslovakia	LV	Latvia	TG	Togo
CZ	Czech Republic	MC	Monaco	TJ	Tajikistan
DE	Germany	MD	Republic of Moldova	TT	Trinidad and Tobago
DK	Denmark	MG	Madagascar	UA	Ukraine
ES	Spain	ML	Mali	US	United States of America
FI	Finland	MN	Mongolia	UZ	Uzbekistan
FR	France			VN	Viet Nam
GA	Gabon				

### Self-sustaining hydrogen generator

This invention concerns improvements in reformers, more especially it concerns improvements in self-sustaining reformers that start-up from ambient  
5 temperature. The volume-specific hydrogen output of these improved reformers makes them suitable for application in fuel-cell powered vehicles.

We have disclosed in EP 0 217 532, a self-igniting partial oxidation reformer or catalytic hydrogen generator which has become known as the "Hot Spot"  
10 reactor. The basic concept is that methanol and air are co-fed into a reactor containing a packed bed of copper on refractory support catalyst, with a down-stream zone containing platinum or palladium catalyst mixed with copper catalyst. The down-stream zone provides self-ignition to raise the reactor temperature to a point

at which a hot spot formed around the point of injection of feedstock into the bed of catalyst. This concept was further developed in the invention of EP 0 262 947, which uses this reactor design to produce hydrogen from hydrocarbons, but using a catalyst composed of platinum and chromium oxide on a support. Further details  
5 have been given in a paper in Platinum Metals Review, 1989, 33, (3) 118-127.

As predicted in the above-mentioned prior art, the use of liquid fuels as hydrogen sources for fuel-cell powered vehicles, or even static systems, is attracting considerable interest. The best established conversion system is steam  
10 reforming, but this is an endothermic reaction, requiring continuous input of energy. Whilst self-sustaining reforming is clearly an interesting concept, in further studies of the Hot-Spot system, we have found that the reactor as described did not permit scale-up, and therefore there was a need to find alternative and improved reactor and/or system designs.

15

The present invention provides a self-igniting and self-sustaining hydrogen generation reactor for a feedstock comprising an organic fuel in liquid, atomised (spray), vapour or gas form and a source of oxygen, such as air, said reactor comprising a permeable fixed bed of copper-supported catalyst and a PGM-  
20 supported catalyst, said reactor further comprising an inlet for the feedstock which comprises a multipoint entry causing high velocity feedstock injection into the bed together with a significant pressure drop and gas expansion, for example an extended surface area of porous ceramic. Preferably, the inlet is in the form of a porous

ceramic tube having a closed end, but other inlet designs may give advantages in particular circumstances.

5 The invention also provides a method for the production of hydrogen from a feedstock comprising an organic fuel and a source of oxygen, such as air, and over a mass of catalyst comprising copper and palladium or other PGM moieties, characterised in that the fuel is in spray, vapour or gas form and further characterised in that the feedstock enters a radial catalyst bed at a multipoint entry which causes a significant pressure drop and results in high velocity injection of feedstock into the  
10 bed. For example the multipoint entry may be by means of an inlet having an extended surface area of porous ceramic.

The porous ceramic used in initial tests was a commercial ceramic of pore size 100µm, purchased from Fairey, England, and this material is therefore  
15 recommended, but we expect that alternative porous ceramic materials will provide substantially similar results. It appears that porous metal inlets, whilst having similar gas flow characteristics to porous ceramics, are unsuitable because heat conduction interferes with the reactions being carried out, and could even cause premature ignition of the methanol/air mixture, for example within the feed tube. However, it  
20 is possible that further material developments, and/or further reactor design improvements, may permit composite metal/ceramic inlet design, which may bring advantages, such as in structural strength.

The organic fuel can be a hydrocarbon or an oxygenate, either in the liquid or gas phase. For ease of description, a liquid oxygenate, methanol, will be referred to in the following description.

5 In the method of the invention, a hot zone forms around the region where the feedstock enters the mass of catalyst. It is believed that this hot zone is preferably at a temperature of 350-600°C. In this zone, methanol is partially oxidised to form CO<sub>2</sub> and H<sub>2</sub>, and some CO. In a preferred embodiment, water is co-fed with the methanol. The presence of water has several beneficial effects:

- 10 (i) It promotes the water-gas shift reaction; thus improving the efficiency of the generator and lowering the concentration of CO in the reformat from 2-8% to 1-5% by volume, and raising concurrently the H<sub>2</sub> concentration.
- (ii) It allows greater thermal control within the catalyst bed.
- (iii) Under some conditions, it can react directly with methanol by steam  
15 reforming, leading to a high yield of H<sub>2</sub>.

The water can be pre-mixed with the methanol feed. The water concentration in the aqueous mixture should be between 1% and 40% by mass, with the preferred concentration being in the range 10-30%.

20

Conventional hydrogen-generators that depend solely or predominantly on the steam-reforming of an organic reactant require continuous input of energy during operation to supply the endothermic heat of reaction. For example, where the organic reactant is methanol:



In the present invention, self-sustaining hydrogen generation is achieved by catalysing the exothermic conversion of the reactant, so obviating the need for continuous input of energy:



We have discovered that Cu-based catalysts are very effective for the latter reaction. It is necessary, however, to add a small amount of precious metal catalyst to provide auto-ignition, and to raise the catalyst bed temperature to a level at which the partial oxidation becomes self-sustaining. When the precious metal catalyst is Pd/SiO<sub>2</sub>, containing 3% palladium by mass, it needs only to be present as 5% of the total mass of the catalyst bed in order to induce ignition from room temperature. Other pre-heating methods may be used, however.

It is believed, although we do not wish to be limited to any expression of theory, that there are several key physical and chemical features in the design of an effective self-sustaining partial-oxidation reformer. The flow restriction caused by the porous ceramic creates a pressure drop, resulting in the pore openings acting as small injection points that supply fluid flow at high velocity into the catalyst bed. Then, in order for the conversion of methanol to occur exclusively by the catalytic partial-oxidation route, both the following conditions must apply within the catalyst bed:

- (a) the reactant velocity must exceed the homogeneous flame speed,

(b) the catalyst must be active enough for the heterogeneous reaction to occur within the short residence time.

To eliminate the possibility of product hydrogen undergoing oxidation, it must be prevented from re-entering a high temperature region in which oxygen is present. To do this, the pressure drop across the ceramic must be high enough to prevent back-flow of products. If the porous ceramic is in the form of a tube or cylinder, the catalyst bed should be arranged concentrically to allow radial flow, and so ensure a minimum path-length through the bed.

It is desirable to reduce or minimise the amount of CO in the product gas fed to the fuel cell. This can be achieved by one or both of passing the product gas over a selective oxidation catalyst that selectively oxidises CO in preference to hydrogen, and passing the product gas over a low temperature water-gas shift catalyst.

The invention will now be illustrated by the following Examples, and with reference to the accompanying drawings which are schematic cross-sections of reactors according to the invention.

### EXAMPLE 1

A batch of catalyst was made by adding copper(II) ethanoate (71g) to aqueous ammonia, which had been prepared by adding distilled water (3150cm<sup>3</sup>) to



a concentrated solution (0.880) of ammonia (350cm<sup>3</sup>). This ammoniacal copper-solution was added to silica spheres (Shell S980B; 750g) in a rotating blender, which was operated for 1 hour to ensure thorough mixing. The silica spheres were then isolated and washed three times with distilled water, before being dried (110°C; 16hr) and calcined (450°C; 2hr). Elemental analysis showed that the resultant copper-loading on the silica spheres was 1.9% (by mass). Before being loaded into hydrogen-generator, the catalyst was activated by reduction (400°C; 2hr; 10% H<sub>2</sub>/N<sub>2</sub>).

A schematic diagram of the hydrogen-generator is shown in Figure 1.

The reactor shell consisted of a stainless cylinder (height 17cm, diameter 9cm), 1, having a main inlet aperture, 2, and an outlet aperture, 3, to which was attached an outlet tube, 4. Fitted through the aperture 2 was a feedstock inlet tube, 5, having a flange, 6. Fitted and sealed to a stainless flange 6 was a porous ceramic tube, 7, held under compression by and sealed to a stainless steel clamping bar, 8. The feedstock inlet tube 5? had an L-shaped manifold, 9, into which was fed an air line, 10, and a feed line for the liquid feed, 11. The liquid-feed line 11 was coiled around the inlet tube 5 for heat exchange, before terminating mid-way down the porous ceramic tube 7. In this embodiment, a further air line, 12, entered the reactor 1 and was coiled around the reactor wall, for preheating the air, before entering the manifold 9. In this configuration, the fluid emerged radially from the porous ceramic tube, but it then followed an axial path through most of the catalyst bed. The volume of the catalyst bed was 850cm<sup>3</sup>.

5 A mixture comprising 85% methanol (by mass) in water was injected (at a rate of  $5.2\text{cm}^3\text{ min}^{-1}$ ) into the catalyst bed through the porous ceramic tube (dimensions: 3cm diameter, 4cm long; pore diameter:  $100\mu\text{m}$ ). The exit stream was cooled to remove any condensable components, and the dry gas was analysed for  $\text{H}_2$ ,  $\text{CO}$ ,  $\text{CO}_2$ ,  $\text{O}_2$  and  $\text{N}_2$ .

10 The bed temperature was raised (by applying an external heat source) until the onset of reaction, which occurred at *ca*  $150^\circ\text{C}$ ; the external heating was then switched off. The composition of the exit stream was recorded as a function of time-elapsed after the onset of reaction (Table 1).

15 A steady state was reached after about 20 minutes of operation, when the hydrogen yield levelled at  $245\text{ litres h}^{-1}$ . When the volume of catalyst was taken into account, the specific output of hydrogen at steady-state was 290 litres per hour per litre of catalyst. At this stage, the maximum temperature within the catalyst bed (in the immediate vicinity of the ceramic tube) stabilised at *ca*  $600^\circ\text{C}$ , whereas the temperature inside the tube remained  $<100^\circ\text{C}$ . Furthermore, the efficiency of the generator was close to 100% (*ie* each molecule of methanol converted produced 2 molecules of hydrogen). When the water was excluded from the reactant-feed, 20 the efficiency declined to 92%.

**TABLE 1**

Output of axial hydrogen-generator with 4cm long ceramic injector  
and 850cm<sup>3</sup> catalyst bed; composition of reformat and hydrogen-yield

Time elapsed

	2 min	15 min	25 min	30 min
H <sub>2</sub> / %	30	31	39	39
CO / %	3	2	2	1.8
CO <sub>2</sub> / %	>20	>20	18	19
O <sub>2</sub> / %	0	0	0	0
N <sub>2</sub> / %	37	37	39	39
H <sub>2</sub> -yield / litres h <sup>-1</sup>	180	195	245	246

**EXAMPLE 2**

The procedure described in Example 1 was followed, except that a longer ceramic injector (dimensions: 3cm diameter, 6cm long) was used and the reactant feed-rates were increased (85% methanol/water 7.6cm<sup>3</sup> min<sup>-1</sup>; air 9000cm<sup>3</sup> min<sup>-1</sup>). After onset of reaction, the generator reached a steady-state after about 15 minutes of operation (Table 2). In this case, the specific hydrogen output at steady state was 420 litres per hour per litre of catalyst.

**TABLE 2**

Output of axial hydrogen-generator with 6cm long ceramic injector and 850cm<sup>3</sup> catalyst bed; composition of reformat and hydrogen-yield.

5

**Time elapsed**

	10 min	15 min	25 min	30 min	40 min
H <sub>2</sub> / %	32	35	36	36	36
CO / %	4	4	4	4.5	5
CO <sub>2</sub> / %	16	16	16	16	16
O <sub>2</sub> / %	0	0	0	0	0
N <sub>2</sub> / %	36	36	37	37	37
H <sub>2</sub> -yield / litres h <sup>-1</sup>	307	358	355	351	362

10

15

**EXAMPLE 3**

The hydrogen-generator shown in Figure 1 was modified, as shown in Figure 2, to allow radial flow through a smaller bed (120cm<sup>3</sup>) of catalyst. As Figure 2 shows, the catalyst bed, 21, was the same length (6cm) as the porous ceramic tube, 22. The bed was enclosed by a copper gauze, 23, which allowed the products to emerge radially. Axial flow through the bed was prevented by the presence of two impervious ceramic plugs, 24 and 25.

20

25

The catalyst was prepared and activated in the way described in Example 1. Neat methanol was injected (at a rate of 3.6cm<sup>3</sup> min<sup>-1</sup>) together with air (2500cm<sup>3</sup> min<sup>-1</sup>) into the catalyst bed through the porous ceramic tube. An external

heat source was applied, but only until the onset of reaction. After onset, a steady state was quickly reached (within 3 minutes), where the specific hydrogen output was 690 litres per hour per litre of catalyst (Table 3). By increasing the air feed-rate (to  $4000\text{cm}^3\text{ min}^{-1}$ ), the hydrogen yield was increased to a maximum of 93 litres  $\text{h}^{-1}$ , which corresponded to a specific output of 775 litres per hour per litre of reactor.

**TABLE 3**

Output of radial hydrogen-generator with 6cm long ceramic injector and  $120\text{cm}^3$  catalyst bed; composition of reformat and hydrogen-yield.

**Time elapsed**

	3 min	25 min	45 min	60 min
$\text{H}_2$ / %	32.5	32	33	32
CO / %	5	5.5	4	3.5
$\text{CO}_2$ / %	16	15.5	16.5	16.5
$\text{O}_2$ / %	0	0	0	0
$\text{N}_2$ / %	43	43	42	42
$\text{H}_2$ -yield / litres $\text{h}^{-1}$	86	80	84	81

**EXAMPLE 4**

A 5%Cu/ $\text{Al}_2\text{O}_3$  catalyst (nominal composition by mass) was prepared by adding copper(II) ethanoate (82.4g) to a concentrated solution (0.880) of ammonia ( $320\text{cm}^3$ ). The resultant solution was added slowly to alumina extrudate (Norton

6173; 500g) in a rotating blender. The extrudate was then heated under nitrogen, using a water bath, until the individual pellets had changed colour (from deep blue to pale blue), indicating that most of the ammonia had been removed. The pellets were subsequently dried (110°C; 16hr), calcined (400°C; 2hr), and activated (400°C; 2hr; 10% $H_2/N_2$ ).

The same preparative steps were also used to make a smaller batch of 5%Pd/ $Al_2O_3$  catalyst, starting from palladium(II) ethanoate (2.64g) and alumina extrudate (same source as above; 23.75g).

10

The 5%Cu/ $Al_2O_3$  and 5%Pd/ $Al_2O_3$  catalysts were mixed together in a ratio of 19:1 by mass. The resulting catalyst mixture was packed as a thin radial bed (2.5mm deep), which fitted concentrically around a porous ceramic tube (11.4cm long; 3cm diameter) and was held in place by a copper gauze (11.4cm long; 3.5cm diameter). The volume of catalyst mixture used in the radial bed was only 55cm<sup>3</sup>.

15

An atomised liquid feed comprising 85% methanol (by mass) in water was injected (at a rate of 3.1cm<sup>3</sup> min<sup>-1</sup>) together with air (3000cm<sup>3</sup> min<sup>-1</sup>) through the porous ceramic tube into the cold (20°C) bed of mixed catalyst. The bed temperature began to rise immediately, and within a minute it was approaching the steady-state maximum (500°C). At steady-state, the average hydrogen yield was 175 litres h<sup>-1</sup> (Table 4), which corresponded to a specific hydrogen output of 3200 litres per hour per litre of catalyst. The radial catalyst bed and the porous ceramic tube contained within it, could be enclosed in a 150cm<sup>3</sup> canister. When the volume

20

COMPARATIVE EXAMPLE 1

The copper/silica catalyst, as prepared and activated in Example 1, was tested in a reactor fitted with a single point-injector (as described by J W Jenkins, EP 0 217 532. Using the same proportions of methanol/water/air as in Examples 1 and 2, hydrogen-yields which were below 10 litres h<sup>-1</sup> could be generated. However, in order to achieve higher yields (up to a maximum of 15 litres h<sup>-1</sup>), a continuous source of external heating (250W) had to be applied.

**CLAIMS**

1. A self-sustaining hydrogen generator reactor for a feedstock comprising an organic fuel in liquid spray, vapour or gas form and a source of oxygen, said reactor comprising a permeable mass of catalyst, and said reactor further comprising an inlet for the feedstock which comprises restricted multipoint entry causing substantial pressure drop, resulting in high velocity injection into the bed and fluid expansion.
2. A reactor according claim 1, wherein the inlet comprises an extended area of porous ceramic with a majority of the external surface area in contact with a bed of catalyst.
3. A reactor according to claim 2, wherein the porous ceramic is in the form of a tube, with a radial catalyst bed held concentrically around it.
4. A reactor according to any one of the preceding claims, wherein the catalyst contains supported copper.
5. A reactor according to any one of the preceding claims, including means to provide pre-heating.
6. A reactor according to claims 1-4 which includes a catalyst that allows start-up from ambient temperature.



of the canister was taken into account, the reactor-specific hydrogen output was 1150 litres per hour per litre of reactor.

**TABLE 4**

5                      Steady-state output of radial hydrogen-generator with 11.4cm long ceramic injector and 55cm<sup>3</sup> catalyst bed, operating at 100% methanol conversion; composition of reformat and hydrogen yield.

**Time elapsed**

10		165 min	190 min	215 min	260 min	290 min	315 min
	H <sub>2</sub> / %	41.7	43.5	41.7	40.4	40.6	40.4
	CO / %	2.9	3.1	3.1	3.1	3.2	3.3
	CO <sub>2</sub> / %	19.5	19.7	19.7	18.3	18.1	18.3
	O <sub>2</sub> / %	0	0	0	0	0	0
15	N <sub>2</sub> / %	35.3	31.6	33.9	36	35.1	35.6
	H <sub>2</sub> -yield / litres h <sup>-1</sup>	174	182	187	165	170	164

20                      When the reactant stoichiometry was changed to a methanol-rich feed (85% methanol/water; 3.5cm<sup>3</sup> min<sup>-1</sup>; air: 2500cm<sup>3</sup> min<sup>-1</sup>), the average hydrogen yield at steady-state rose to 200 litres h<sup>-1</sup> (Table 5). At the same time, the bed temperature decreased to 370°C, and the methanol conversion dropped from 100% to 70%. However, the hydrogen concentration in the product stream was higher than expected

(on the basis of all the converted methanol undergoing partial oxidation), indicating that some of the hydrogen was being generated by steam reforming.

Under the methanol-rich conditions, the specific hydrogen output was 3600 litres per hour per litre of catalyst, or 1300 litres per hour per litre of reactor when the volume of the canister was taken into account.

**TABLE 5**

Steady-state output of radial hydrogen-generator with 11.4 long ceramic injector and 55cm<sup>3</sup> catalyst bed, operating at 70% methanol conversion; composition of reformat and hydrogen yield.

**Time elapsed**

	65 min	225 min	315 min	410 min	560 min	1050 min
H <sub>2</sub> / %	48.1	47.0	45.9	46.9	47.0	46.8
CO / %	2.3	2.4	2.8	2.7	2.6	2.8
CO <sub>2</sub> / %	19.6	18.3	19.2	19.5	19.2	18.2
O <sub>2</sub> / %	0	0	0	0	0	0
N <sub>2</sub> / %	27	28	29.7	28	27.2	27.3
H <sub>2</sub> -yield / litres h <sup>-1</sup>	192	175	206	219	200	198

7. A method for the self-starting, self-sustaining production of hydrogen from a feedstock comprising a fuel and a source of oxygen and over a mass of catalyst comprising copper and precious metal moieties, characterised in that the feedstock enters the mass of catalyst at a multipoint entry causing a significant pressure drop with high velocity injection of feedstock into mass of catalyst.

8. A method according to claim 7, wherein a hot zone exists in the catalyst mass close to the injection point.

9. A method according to claim 8, wherein the hot zone is at a temperature of approximately 150-600°C.

10. A method according to any one of claims 7 to 9, wherein there is relatively no back flow of product hydrogen.

11. A method according to any one of claims 7 to 10, wherein water is simultaneously injected into the catalyst mass with the feedstock.

12. A method according to claim 11, wherein the added water reacts with some of the methanol to generate hydrogen.

13. A method according to any one of claims 7 to 12, wherein the CO concentration of the product gas is reduced by passing said product gas over a low temperature water-gas shift catalyst.

14. A method according to any one of claims 7 to 12, wherein the CO concentration of the product gas is reduced by passing said product gas together with an added source of oxygen over a catalyst that selectively oxidises CO in preference to hydrogen.

5

15. A method according to claims 7 to 12, wherein the CO concentration of the product gas is reduced by passing said product gas over a low temperature water-gas shift catalyst and then a selective oxidation catalyst in succession.

10

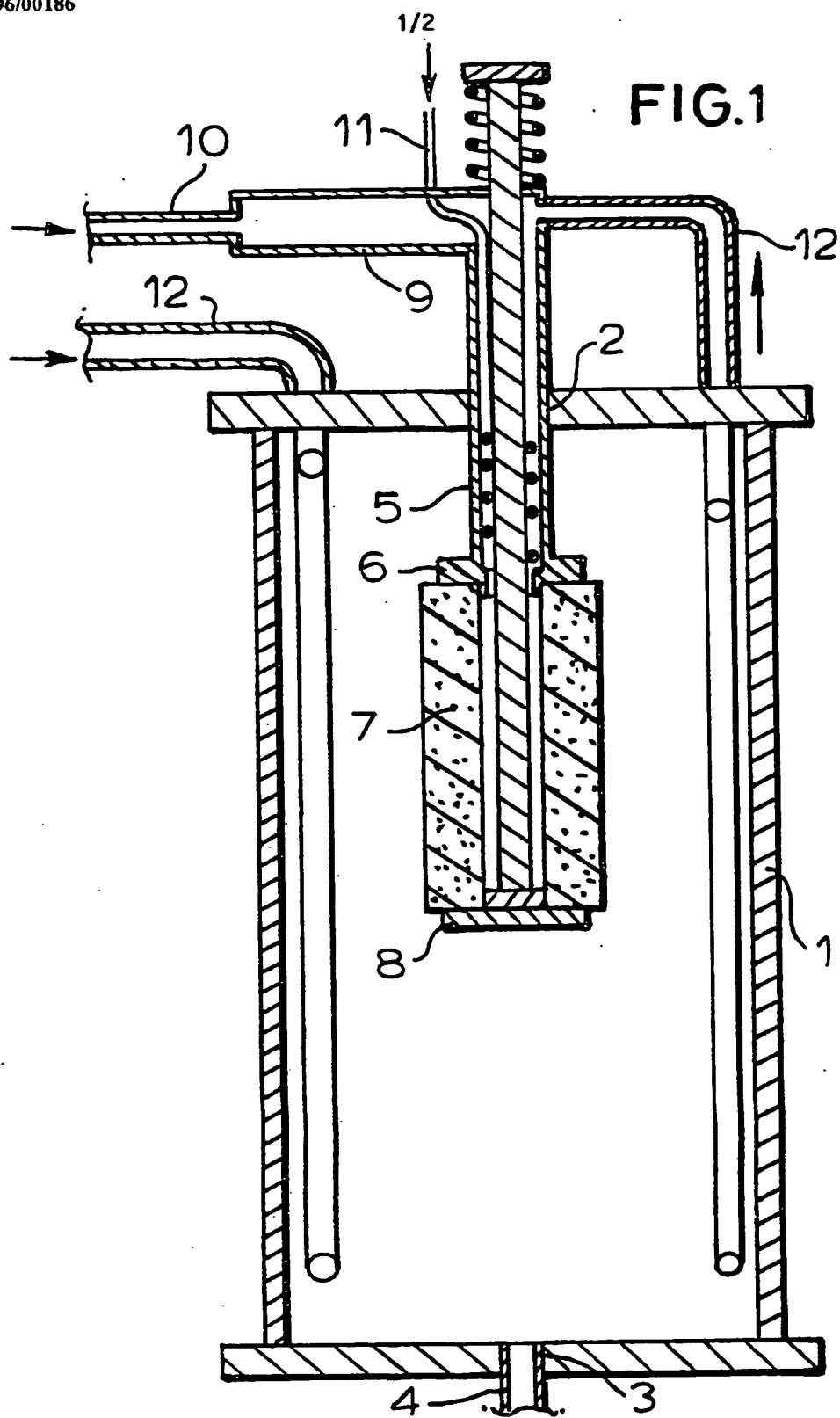
16. A method according to any one of claims 7 to 15, wherein the fuel is methanol.

17. A fuel cell system comprising a hydrogen generator reactor according to any one of claims 1 to 6.

15

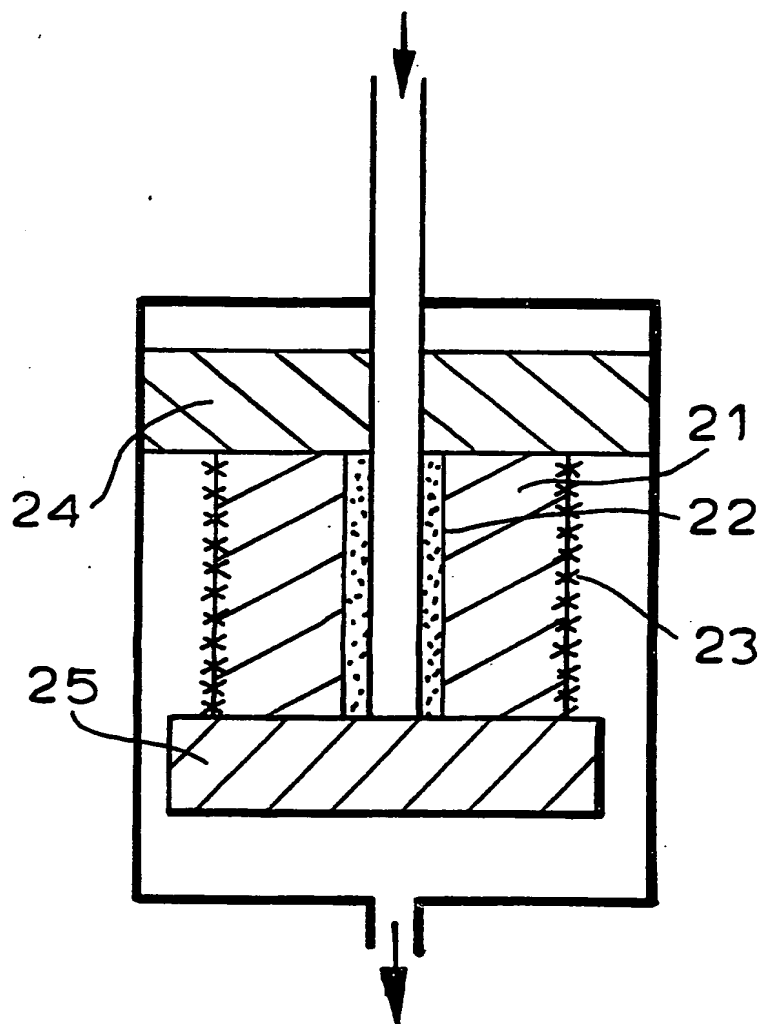
18. A system according to claim 17, in the form of a hydrogen generator according to any one of claims 1 to 6, and a fuel cell producing electricity, for automotive use.

20



SUBSTITUTE SHEET (RULE 26)

FIG. 2



1  
INTERNATIONAL SEARCH REPORT

International application No.

PCT/GB 95/01500

## A. CLASSIFICATION OF SUBJECT MATTER

IPC6: C01B 3/38, B01J 23/72, H01M 8/06

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC6: C01B, B01J, H01M

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

EPOQUE, WPI

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	EP, A1, 0262947 (JOHNSON MATTHEY PUBLIC LIMITED COMPANY), 6 April 1988 (06.04.88), page 1, line 27 - line 48; page 1, line 58 - line 61, figure 1, abstract, example 1 --	1-3,7
A	EP, A1, 0217532 (JOHNSON MATTHEY PUBLIC LIMITED COMPANY), 8 April 1987 (08.04.87), column 1, line 25 - line 50, figure 1, abstract --	1-5,7-9
A	DE, 2220617 (METALLGESELLSCHAFT AG), 15 November 1973 (15.11.73), claims 1,2,5 -- -----	4-5,9,11-12

☐ Further documents are listed in the continuation of Box C. ☒ See patent family annex.

## \* Special categories of cited documents

"A" document defining the general state of the art which is not considered to be of particular relevance

"E" earlier document but published on or after the international filing date

"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance: the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

"Y" document of particular relevance: the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art

"&amp;" document member of the same patent family

Date of the actual completion of the international search

3 October 1995

Date of mailing of the international search report

20. 10. 95

Name and mailing address of the International Searching Authority

European Patent Office, P.B. 5818 Patentlaan  
NL-2280 HV Rijswijk  
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,  
Fax: (+31-70) 340-3016

Authorized officer

Igor Schaber  
Telephone No.

**INTERNATIONAL SEARCH REPORT**  
Information on patent family members

28/08/95

International application No.

PCT/GB 95/01500

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP-A1- 0262947	06/04/88	AU-B- 603832	29/11/90
		AU-A- 7899787	14/04/88
		DE-A- 3779543	09/07/92
		JP-A- 63147802	20/06/88
		US-A- 4897253	30/01/90
EP-A1- 0217532	08/04/87	AU-A- 6215686	05/03/87
		JP-A- 62059501	16/03/87
		US-A- 4789540	06/12/88
DE- - 2220617	15/11/73	BE-A,A- 798785	26/10/73
		FR-A- 2181795	07/12/73
		NL-A- 7303784	30/10/73